

PRACTICE SET A--IONIC COMPOUNDS

Write the formula for each of the following compounds:

iron(II)chloride

sodium iodide

barium nitride

magnesium oxalate

zinc chromate

calcium chlorite

chromium(III)oxide

aluminum fluoride

tin(IV)oxide

copper(II)nitrate

lithium chloride

aluminum sulfide

Write the names for each of the following compounds:

K_2SO_3

$AlBr_3$

Mg_3N_2

$Al_2(SO_4)_3$

$HgCl_2$

$AgNO_3$

$Fe(NO_2)_2$

$Ca_3(PO_4)_2$

$PbBr_2$

Na_2CO_3

$Ca(C_2H_3O_2)_2$

$Mn(OH)_2$

ANSWERS TO PRACTICE SET A--IONIC COMPOUNDS

You can use these answers as another worksheet.

$FeCl_2$

NaI

Ba_3N_2

MgC_2O_4

$ZnCrO_4$

$Ca(ClO_2)_2$

Cr_2O_3

AlF_3

SnO_2

$Cu(NO_3)_2$

$LiCl$

Al_2S_3

potassium sulfite

aluminum bromide

magnesium nitride

aluminum sulfate

mercury(II)chloride

silver nitrate

iron(II)nitrite

calcium phosphate

lead(II)bromide

sodium carbonate

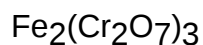
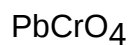
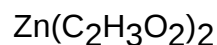
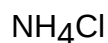
calcium acetate

manganese(II)hydroxide

PRACTICE SET B FOR IONIC AND COVALENT COMPOUNDS

Identify the following compounds as ionic or covalent (molecular)

Write the names of the following:



Write the formula for the following compounds

barium bromide

iron(II)oxide

sodium phosphate

nickel(II) chlorate

dinitrogen tetrachloride

calcium sulfide

copper(II)chloride

barium sulfate

cobalt(II)ntrate

sodium iodide

cobalt(II)sulfide

calcium acetate

carbon monoxide

chromium(III)sulfate

phosphorous trichloride

magnesium nitride

aluminum sulfite

phosphorous pentachloride

(Answers are given on the next page)

Answer Sheet for Practice Set B for Ionic and Covalent Compounds

sodium sulfide
ionic

silver chloride
ionic

aluminum bromide
ionic

diphosphorous pentoxide
covalent

potassium nitride
ionic

aluminum sulfate
ionic

iron(II)oxide
ionic

dinitrogen trioxide
covalent

copper(I) sulfide
ionic

ammonium chloride
ionic

iron(III)oxide
ionic

carbon tetrachloride
covalent

barium nitrate
ionic

copper(I) nitrate
ionic

zinc acetate
ionic

lead(II)chromate
ionic

tin(IV)sulfite
ionic

iron(III)dichromate
ionic

BaBr₂
ionic

FeO
ionic

Na₃PO₄
ionic

Ni(ClO₃)₂
ionic

N₂Cl₄
covalent

CaS
ionic

CuCl₂
ionic

BaSO₄
ionic

Co(NO₃)₂
ionic

NaI
ionic

CoS
ionic

Ca(C₂H₃O₂)₂
ionic

C O
covalent

Cr₂(SO₄)₃
ionic

PCl₃
covalent

Mg₃N₂
ionic

Al₂(SO₃)₃
ionic

PCl₅
covalent

You can use the answer sheet as a worksheet. That is, work backwards and see if you can write the formulas and the names.

Practice Set C--Ionic and Covalent Compounds

Write the name and formula for each of the following:

CaS	sulfur difluoride
Al ₂ O ₃	tin(II) chloride
N ₂ O ₅	germanium dioxide
KClO ₃	ammonium acetate
NaHCO ₃	iron(III) hydrogen sulfate
Rb ₂ O	calcium nitrite
Fe(ClO ₄) ₃	platinum(IV) chloride
Ca ₃ (PO ₄) ₂	chromium (II) hypochlorite
P ₂ O ₅	methane
KMnO ₄	barium sulfide
AuCl ₃	gallium oxide
Pb(NO ₃) ₂	calcium hydroxide
NH ₃	lithium nitride
NaNO ₂	nitrogen trifluoride

Answers Practice Set C

Calcium sulfide	SF ₂
Aluminum oxide	SnCl ₂
Dinitrogen pentaoxide	GeO ₂
Potassium chlorate	NH ₄ C ₂ H ₃ O ₂
Sodium hydrogen carbonate (sodium bicarbonate)	Fe(HSO ₄) ₃
Rubidium oxide	Ca(NO ₂) ₂
Iron(III)perchlorate	PtCl ₄
Calcium phosphate	Cr(ClO) ₂
Diphosphorus pentaoxide	CH ₄
Potassium permanganate	BaS
Gold(III) chloride	Ga ₂ O ₃
Lead (II) nitrate	Ca(OH) ₂
Ammonia	Li ₃ N
Sodium nitrite	NF ₃